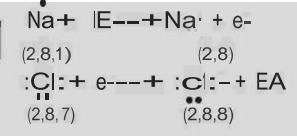


Lithium	Li
Potassium	K
Barium	Ba
Sodium	Na
Calcium	Ca
Magnesium	Mg
Aluminium	Al
Zinc	Zn
Iron	Fe
Nickel	Ni
Tin	Sn
Lead	Pb
Hydrogen	H
Copper	Cu
Mercury	Hg
Silver	Ag
Gold	Au
Platinum	Pt



Electrostatic force of attraction due to which positively & negatively charged ions are bonded with each other is known as ionic bond.

Compound containing ionic bonds is known as ionic compound e.g. NaCl, MgO etc.



Example :
formation of NaCl

Definition

Definition

Properties

- (i) Consist of ions
- (ii) Solid & hard due to strong electrostatic force of attraction
- (iii) Definite crystal structure
- (iv) High m.p. & b.p.
- (v) Soluble in polar solvents like H₂O
- (vi) Brittle nature
- (vii) Good conductor of electricity in molten and aqueous state



Reactivity series

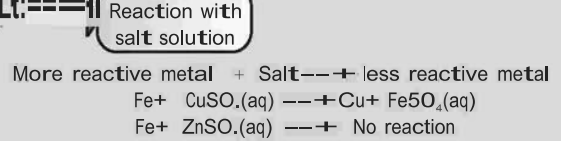
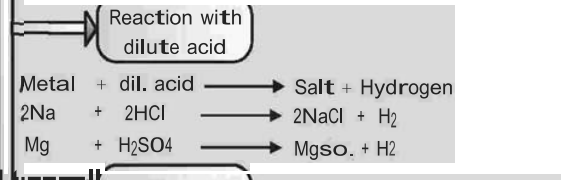
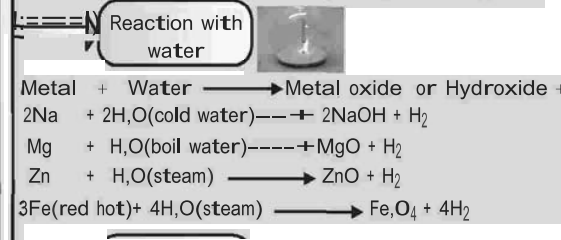
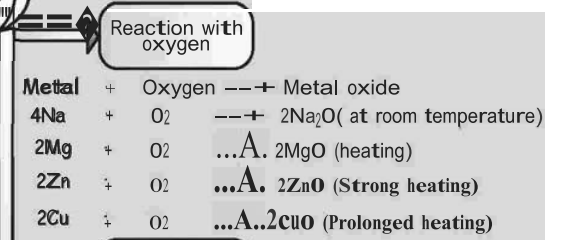
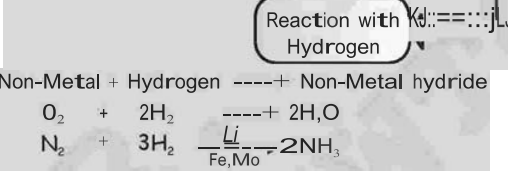
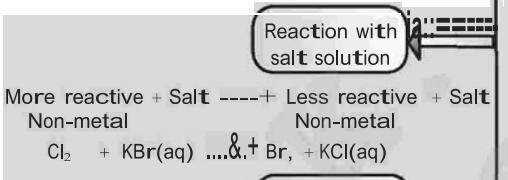
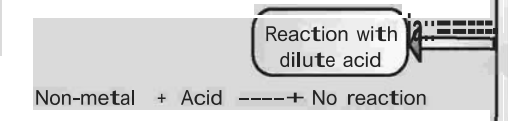
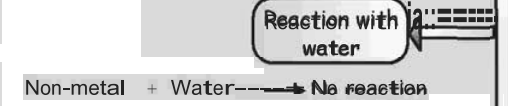
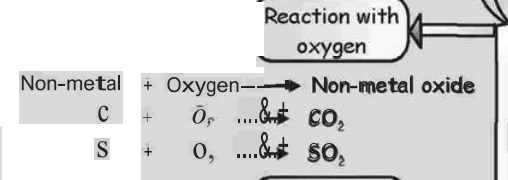
Chemical properties of non-metal

Chemical properties of metal

Physical properties of metal & non-metal

Non-Metal

Metal



- Generally gaseous
- Low m.p. & b.p.
- Brittle
- Generally soft
- Bad conductor of heat & electricity
- Not ductile
- Do not have any lustre
- Not sonorous
- Low density

- Generally Solid
- High m.p. & b.p.
- Malleable
- Generally hard
- Good conductor of heat & electricity
- ductile
- lustrous
- Sonorous
- High density

