

Atomic  
Electro  
Ionic



Ionization energy decreases  
Electronegativity decreases



Ionisation energy increases

Electronegativity increases

### Atomic Radius

Along period, extra orbit is added therefore size of atom increases.  
Along group, no. of protons increases therefore size decreases.

Largest atom **Cs**

### Electropositivity

Ability of a metal to lose electrons increases with increase in its size.

Largest electropositive atom **Cs**

### Ionic Size

Ionic size depends on  $\frac{e}{p}$  ratio.  
As protons increase, attraction over electron increases therefore size of ion decreases.

### Ionisation Energy

As size increases it becomes easy to remove outer electron. Thus ionisation energy decreases along the group

Largest ionisation energy required **F**

### Electronegativity

Ability of an atom to accept electron increases with decrease in its size.

Largest electronegative atom **F**